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Name Class Date Chapter Test A,
continued Use this figure to answer
questions 7 and 8. ____ 7. A solution
containing 35 g of Li_2SO_4 dissolved in
100 g of water is heated from 10°C to
 90°C . According to information in the
figure, this temperature change would
result in a. an additional 5 g of Li_2SO_4
in ...

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Modern Chemistry 27 Chapter Test
Name Class Date Chapter Test A,
continued ____ 4. The electron
configuration below violates a. the Pauli
exclusion principle. b. the Aufbau
principle. c. Hund's rule. d. Both (a) and
(c) ____ 5. A photon is emitted from a
gaseous atom when an electron moves
to its ground state from a(n) a. inner
shell. b ...

Assessment Chapter Test A

Modern Chemistry 86 Chapter Test
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continued ____13. The separation process of paper chromatography can be explained by a. vaporization of the ink. b. water vapor pressure. c. capillary action. d. pull of gravity. ____14. When there is a small decrease in temperature, the average kinetic energy of the particles ...

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continued ____7. The melting points of ionic compounds are higher than the melting points of molecular compounds because a. ionic substances tend to vaporize at room temperature. b. ionic substances are brittle. c. attractive

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forces between ions are greater than the attractive ...

Modern Chemistry Chapter 5 Test

CHEMISTRY CHAPTER 1 REVIEW - Holt
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MATTER AND CHANGE Include graphic
organizer(s) for this chapter LOOK AT
MIDTERM EXAM REVIEW AND ADD
NOTES FROM IT TO ALL CHAPTER
REVIEW PAGES The following pages
contain the bulk (but not all) of the
information for the chapter 1 test.

Modern Chemistry Chapter 1 Review Answers Matter And Change

Modern Chemistry99Chapter Test.
Chapter: Gases. PART I In the space
provided, write the letter of the term or
phrase that best completes each
statement or best answers each
question. _____ 1. If the temperature of a
gas remains constant, then the pressure
of the gas will increase if the. a.mass of
the gas molecules decreases. b.diffusion

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of the gas molecules increases.

Assessment Chapter Test B

MODERN CHEMISTRY The atomic number of nickel-60 is 28. How many neutrons does this iso- tope have? Carbon-14 has 8 neutrons. What is the atomic number of carbon-14? 11
CHAPTER 3 TEST HR W material
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Modern Chemistry 132 Chapter Test
Name Class Date Chapter Test B,
continued Refer to the following
statement to answer questions 30–32:
Dilute $\text{HCl}(\text{aq})$ and $\text{NaOH}(\text{aq})$ are mixed
in chemically equivalent quantities. 30.
Write the chemical equation for the
reaction. 31. Write the overall ionic
equation for the reaction. 32. Write the
net ionic equation.

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Modern Chemistry Chapter 7 Test Answers

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CHAPTER 13 REVIEW Ions in Aqueous Solutions and Colligative Properties

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MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Match the four compounds on the right to their descriptions on the left.

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