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reactants, just as the reactants can react to give the products. yes.

Disturbing Equilibrium Lab Flashcards | Quizlet

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Sem 2 Points Possible: 50 Name: Sana
Hammad Date: 10/16/2019 Answer

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these questions before beginning the lab.

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It is possible, however, to disturb that equilibrium by changing conditions. For example, you could increase the concentration of one of the products, or decrease the concentration of one of the

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reactants, or change the temperature. When a change of this type is made in a reaction at equilibrium, the reaction is no longer in equilibrium.

15.12: Disturbing a Reaction at Equilibrium- Le Châtelier ...

Disturbing Equilibrium Lab Answer Key
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Equilibrium Lab Report Title: Equilibrium
Lab Report Objective(s): To observe the
changes in equilibrium when different
components are added or taken away.
Hypothesis: Adding reactants will shift

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the equilibrium to the right, and taking away reactants will shift the equilibrium to the left. Procedure: Controlled variables: The amount of substance in each starting test tube, the size of each ...

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The lab taught the concepts of torques, and rotational equilibrium. Introduction Equilibrium is a goal of the builders of many different structures. In physics an object is said to be in mechanical equilibrium if it is in a state of translational and rotational equilibrium. Torque Lab.docx - Google Docs

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Disturbing Equilibrium Lab 38 Answers

explain, using Le Châtelier's principle, the equilibrium shift you observed. 4. What effect would a catalyst have on the relative amounts of reactants and products in this reaction at equilibrium? 5. This lab dealt with a chemical equilibrium. Give an example of a

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physical equilibrium. 6.

LE CHATELIER'S PRINCIPLE EXPERIMENT 20

$37 [\text{CoCl}_4]^{2-}(\text{alc}) + 6 \text{H}_2\text{O}(\text{l}) \rightleftharpoons$
 $[\text{Co}(\text{H}_2\text{O})_6]^{2+}(\text{alc}) + 4 \text{Cl}^-(\text{alc})$ Eq 2 The
[CoCl₄]²⁻-complex ion produces a
brilliant blue color while the [Co(H₂O)₆]²⁺
ion produces a rose- pink color.

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Again by observing changes in the color and by monitoring the absorbance of the system resulting from placing a stress on the system we can monitor the equilibrium of the system.

Experiment 6: Equilibrium and Le Châtelier's Principle

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this is my chemistry lab
discussion&conclusion please help me
revise it (mainly focus on conclusion if
possible) thank you! Discussion In the
equilibrium experiment we were able to
observe the direction of the equilibrium
by looking at our chemical equation
 $\text{CH}_3\text{COOH} + \text{C}_2\text{H}_5\text{OH} \rightleftharpoons \text{CH}_3\text{COOC}_2\text{H}_5$

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Answer to Disturbing Equilibrium LAB 7

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Questions for Disturbing Equilibrium
Table 7.1 Well number, contents Color
A1, indicator+(1...

Solved: Disturbing Equilibrium LAB 7 Questions For Disturb ...

Back in the late 1800s, the French chemist Henri Le Chtelier came up with a rule about equilibria that we still use

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extensively today, to the chagrin of many chemistry students. Le Chtelier's Principle states that if you change the conditions of an equilibrium, the equilibrium will shift in a way ...

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