

Experiment 3 Ester Formation Preparation Of Benzocaine

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Experiment 3 Ester Formation Preparation

Topic: Experiment 3 page 1 of 3. Experiment 3. Ester Formation: Preparation of Benzocaine. Esterification [see Ege's: Sections 15.6 B-D, pp 622-626] $\text{H}_3\text{C}-\text{CO}-\text{OH} + \text{C}_2\text{H}_5\text{OH} \rightleftharpoons \text{H}_3\text{C}-\text{CO}-\text{OC}_2\text{H}_5 + \text{H}_2\text{O}$ The experimental equilibrium constant for the reaction above is: $K_{\text{eq}} = \frac{[\text{ethyl acetate}] [\text{H}_2\text{O}]}{[\text{acetic acid}] [\text{ethanol}]} = 3.38$

Experiment 3. Ester Formation: Preparation of Benzocaine.

Experiment 3: Enolate Alkylation, Ester Hydrolysis, and Amide Bond Formation 3 You will then synthesize a mixture of the diastereomeric amides R,R-9 and S,R-9 by reaction of racemic 2-phenylpropanoic acid (8) with the optically pure amine, (R)-Ph(Me)CHNH₂. These diastereomeric amides will then be separated by MPLC.

Experiment 3 Enolate Alkylation, Ester Saponification, and ...

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Sodium carbonate solution, Na_2CO_3 (aq) – see CLEAPSS Hazard HC095A and CLEAPSS Recipe Book RB080. Procedure. Add 10 drops of ethanoic acid (or propanoic acid) to the sulfuric acid in the specimen tube. Add 10 drops of ethanol (or other alcohol) to the mixture. Put about 10 cm³ of water into the 100 cm³ beaker. Carefully lower the tube into the beaker so that it stands upright.

Making esters from alcohols and acids | Experiment | RSC

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5 cm³ ethanol + 6 cm³ butanoic acid. Add 4 drops of concentrated sulphuric acid to each test tube. Pour about 150 ml of tap water in a 250 ml beaker. Place the test tubes in the water and heat the water on a hot plate to a temperature of about 60°C – 75°. Leave the test tubes in the hot water bath for 15 minutes.

Preparation and identification of Esters - Physical ...

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The current experiment uses the carboxylic acid derivative, acetic anhydride, for ester formation. The advantage of using acetic anhydride is that you do not produce water which can be used for hydrolysis of the newly formed ester. Concentrated sulfuric acid will be used to keep everything in the protonated state.

Chemistry 211 Experiment 3 - MiraCosta College

To make a small ester like ethyl ethanoate, you can gently heat a mixture of ethanoic acid and ethanol in the presence of concentrated sulphuric acid, and distil off the ester as soon as it is formed.

Preparation of Esters - Chemistry LibreTexts

PROCEDURE. 1. Fill a 250 mL beaker half full with water, and heat the water to 60°C with a Bunsen burner. 2. To a small test

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tube, add approximately 1 g of a salicylic acid and 1 mL of methanol. 3. Add 3–5 drops of concentrated sulfuric acid to the tube and heat the tube in a water bath at 60°C for 15 minutes. 4.

PREPARATION OF AN ESTER EXPERIMENT 27

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For the synthesis of each ester, add 20 drops of the alcohol and 10 drops of the carboxylic acid (or enough solid on the end of a spatula that is about the size of a large match head) to a small, labeled test tube. ⚠Use extra care when working with the concentrated sulfuric acid (H. 2. SO.

Experiment 13 What's That Smell? (Synthesis of Esters)

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Grade 12 Practical Investigation Preparation Of An Ester

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The reaction we wish to carry out is synthesis of a fragrant ester via acid-catalyzed esterification. For example, the reaction for butyl butyrate is: $\text{OH} + \text{H} + \text{O} + 2\text{O} \text{HO O O}$ The purpose of this experiment is to illustrate one way to make a n ester from a carboxylic acid and an alcohol . Activation by acid has been chosen for th ree reasons: 1.

SYNTHESIS OF AN ESTER - Linfield University

SCH4C Lab #13 - Preparation and Purification of an Ester. Introduction: The reaction between a carboxylic acid with an

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alcohol results in the formation of an ester. The reaction is carried out in a concentrated acid, such as sulphuric to help remove the water from between the acid and alcohol

Lab Preparation and Purification of an Ester

This is an annotated practical demonstration covering how an ester can be prepared from an alcohol and a carboxylic acid. As new videos are published, they w...

Part 1: Synthesis of an Ester from Alcohol and Carboxylic

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When butter becomes rancid, the odor is caused by the formation of butyric acid (butanoic acid), $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$, the same compound that is responsible for the odor of sweat. When a carboxylic acid reacts with an alcohol the product is often a sweet smelling compound called an ester.

ESTERS: THE PREPARATION AND IDENTIFICATION OF AN

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Experiment 3: Preparation of Synthetic Scents and Flavors Introduction Esters, were a type of organic compound that were responsible for the fragrance and flavor of flowers and fruits. For example were the isopentyl acetate in bananas, methyl salicylate in wintergreen, and ethyl butyrate in pineapples.

Laboratory Report on Preparation of Esters.docx ...

Science Process Skills 2:11:5.5 Students will discover how the composition of a molecule affects its interactions with other molecules. Science Process Skills 4:12:2.2 Students will create written reports and journals to share and communicate

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