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## Chapter 20 Redox Reactions Answers

Chapter 20 Worksheet: Redox ANSWERS I. Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.

- $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$   
Sr 0 to  $\text{Sr}^{2+}$ ; oxidized/reducing agent  $\text{O}_0$  to  $\text{O}^{2-}$ ; reduced/ox. ag.
- $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$  Li 0 to  $\text{Li}^{1+}$ ; oxidized/red. ag.  $\text{S}_0$  to  $\text{S}^{2-}$ ; reduced/ox. ag.
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## **Chapter 20 - Oxidation-Reduction Reactions - 20.2 ...**

636Chapter 20Redox Reactions Figure 20-1. The reaction of magnesium and oxygen involves a transfer of electrons from magnesium to oxygen. Therefore, this reaction is an oxidation-reduction reaction. Using the classifications given in Chapter 10, this redox reaction also is classified as a combustion reaction.

## **Chapter 20: Redox Reactions**

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Chapter 20 Worksheet: Redox ANSWERS I. Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.

1.  $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$   
Sr 0 to  $\text{Sr}^{2+}$ ; oxidized/reducing agent  $\text{O}_2$  to  $\text{O}^{2-}$ ; reduced/ox. ag.
2.  $2\text{Li} + \text{S} \rightarrow 2\text{Li}_2\text{S}$   
Li 0 to  $\text{Li}^{+}$ ; oxidized/red. ag. S 0 to  $\text{S}^{2-}$ ; reduced/ox. ag.
3.  $2\text{Cs} + \text{Br}_2 \rightarrow 2\text{CsBr}$

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In Chapter 19, you read that all redox reactions involve a transfer of electrons from the species that is oxidized to the species that is reduced. Figure 20.1 and Figure 20.2 illustrate the simple redox reaction in which zinc atoms are oxidized to form zinc ( $\text{Zn}^{2+}$ ) ions. The two elec-

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## **Chapter 20: Electrochemistry**

20.1: Oxidation States & Redox Reactions Oxidation state is a useful tool for keeping track of electron transfers. It is most commonly used in dealing with metals and especially with transition metals. Oxidation signifies a loss of electrons and reduction signifies a gain of electrons.

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Prentice Hall Chemistry, chapter 20 "oxidation-reduction reactions" Terms in this set (49) True or false: oxidation and reduction always occur simultaneously. True. The substance gaining oxygen is \_\_\_\_\_ ... An equation showing just the oxidation or just the reduction that takes place in a redox reaction.

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cell an electrochemical cell in which the redox reaction occurs naturally and produces electrical energy. electrode potential. ... Chapter 20- Electrochemistry. 22 terms. Chapter 14 Acid and

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Bases. 36 terms. Chemistry Chapter 17 - Chemical Thermodynamics. 11 terms.

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Chapter 20: Redox Reactions  
trons also is an oxidation-reduction reaction  
636 Chapter 20 Redox Reactions Figure 20-1  
The reaction of magnesium and oxygen involves a transfer of electrons from magnesium to oxygen  
Therefore, this reaction is an oxidation-reduction reaction  
Using

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This overall (or net) reaction is the following redox reaction:  
$$\text{Zn(s)} + \text{Cu}^{2+}(\text{aq}) \rightarrow \text{Zn}^{2+}(\text{aq}) + \text{Cu(s)}$$
  
Although the two half-reactions occur at the same time, they occur at different places in the cell. Thus, for the reaction to proceed, electrons must pass through the wire that connects the two half-cells.



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## **CHAPTER 20 Electrochemistry**

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## **Chapter 20 - Oxidation-Reduction Reactions - Standardized ...**

636 Chapter 20 Redox Reactions Figure 20-1 The reaction of magnesium and oxygen involves a transfer of electrons from magnesium to oxygen Therefore, this reaction is an oxidation-reduction reaction Using the classifications given in Chapter 10, this redox reaction also is classified as a combustion reaction

$$2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$$

Each

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Chapter 20 "Oxidation-Reduction Reactions". Use these reactions to learn the vocabulary and major concepts presented in this chapter. a process that involves a complete or partial gain of electrons or the loss of oxygen; it results in a decrease in the oxidation number of an atom.

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